

Q - How many translational, rotational and vibrational degree of freedoms are shown by He, H_2 , HCl, CO_2 , H_2O and CH_4 .

Ans -

① He - No. of atoms, $n = 1$ \therefore No. of degree of freedom $= 3n = 3$
 $\therefore N_{tr} = 3$, $N_{rot} = 0$, $N_{vib} = 0$

② H_2 - No. of atoms, $n = 2$
No. of degree of freedom $= 3n = 3 \times 2 = 6$
 $\therefore N_{tr} = 3$, $N_{rot} = 2$, $N_{vib} = 3 \times 2 - 5 = 1$

HCl - No. of atoms, $n = 2$
No. of degree of freedom $= 3n = 3 \times 2 = 6$
 $\therefore N_{tr} = 3$, $N_{rot} = 2$, $N_{vib} = 3 \times 2 - 5 = 1$

Teacher's Signature

④ CO_2 - No. of atoms, $n = 3$, A Linear molecule.
 No. of degree of freedom $= 3n = 3 \times 3 = 9$
 $\therefore N_{\text{tr}} = 3$, $N_{\text{rot}} = 2$; $N_{\text{vib}} = 3 \times 3 - 5 = 4$

⑤ H_2O , No. of atoms, $n = 3$, A nonlinear molecule.
 No. of degree of freedom $= 3n = 3 \times 3 = 9$
 $\therefore N_{\text{tr}} = 3$, $N_{\text{rot}} = 3$, $N_{\text{vib}} = 3 \times 3 - 6 = 3$

⑥ C_2H_2 , No. of atoms, $n = 12$, A nonlinear molecule.
 No. of degree of freedom $= 3n = 3 \times 12 = 36$
 $\therefore N_{\text{tr}} = 3$, $N_{\text{rot}} = 3$, $N_{\text{vib}} = 3 \times 12 - 6 = 30$

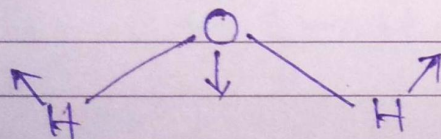
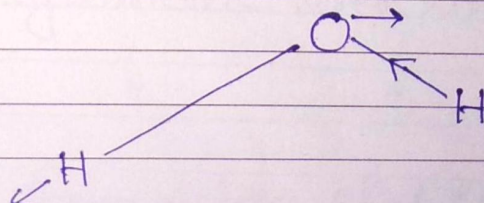
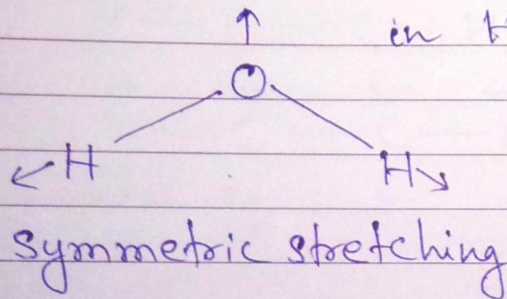
Q - Diagrammatically represents the different normal modes of vibrations of CO_2 , H_2O and C_2H_2 .

Soln

① H_2O - $n = 3$, A non-linear molecule

$$\therefore N_{\text{vib}} = 3n - 6 = 3 \times 3 - 6 = 3$$

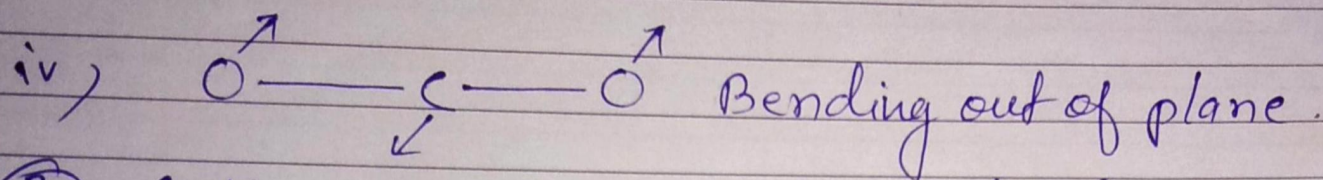
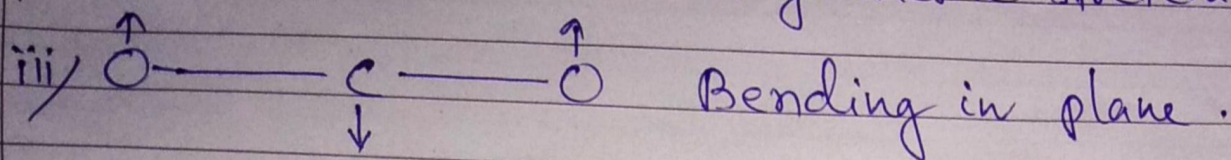
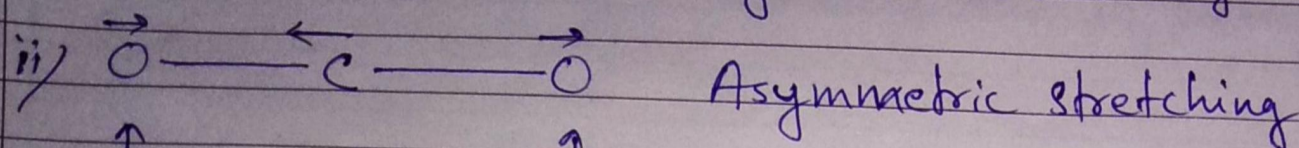
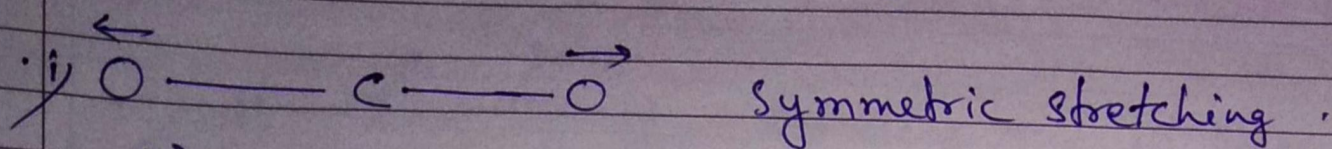
There are three normal modes of vibration in H_2O .



2) CO_2 - $n=3$, A linear molecule

$$N_{\text{vib}} = 3n - 5 = 3 \times 3 - 5 = 4$$

Four normal modes of vibrations are possible.



3) C_2H_2 - $n=4$, A linear molecule.

$$N_{\text{vib}} = 3n - 5 = 3 \times 4 - 5 = 7$$

