

EXPERIMENT No. ~ 06NAME OF THE EXPERIMENT :-

Preparation of alum.

APPARATUS REQUIRED :-

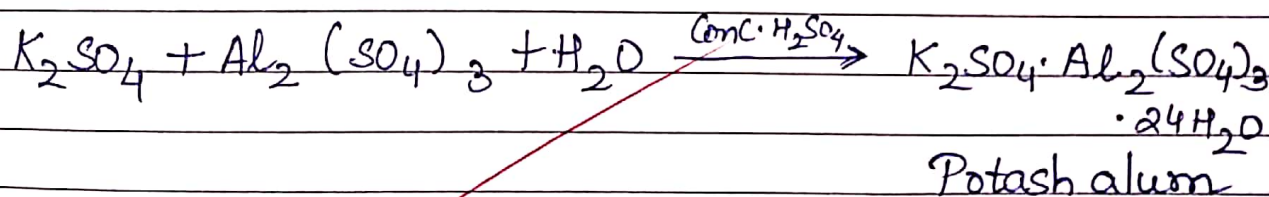
Basin, Beaker, Test tubes, Test tube stand, Test tube holder, Wire Gauge, Tripod Stand, watch glass.

CHEMICALS REQUIRED :-

Aluminium Sulphate, Concentrated sulphuric acid, Potassium Sulphate.

THEORY :-

When equimolecular proportion of potassium sulphate and aluminium sulphate solution are mixed in presence of few drops of concentrated sulphuric acid ( $H_2SO_4$ ), ~~due to the~~ potash alum is obtained.

Reaction

Teacher's Signature : \_\_\_\_\_

PROCEDURE :-

In a beaker, 6.7 g of aluminium sulphate was dissolved in 50 ml of water and 2-3 ml of conc.  $H_2SO_4$  was added in it. Also, in another beaker 1.7 g of  $K_2SO_4$  was dissolved in minimum volume of water. Then, either of the two solution was warmed and mixed together in a basin and allowed to evaporate for crystallization to get beautiful crystal of alum.

RESULT :-

The yield of alum is about 9.455 g.